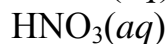
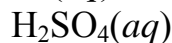
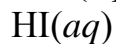
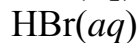
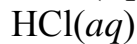
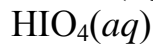
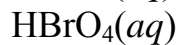
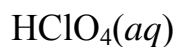
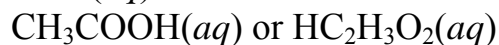
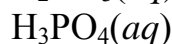
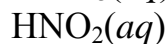
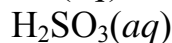
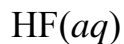


EXAMPLES OF ACIDS AND BASES

Common Strong Acids:



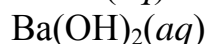
Common Weak Acids:



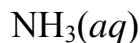
Other organic acids (such as citric acid)

Common Strong Bases:

All alkali metal hydroxides: $\text{LiOH}(aq)$, $\text{NaOH}(aq)$, $\text{KOH}(aq)$, $\text{RbOH}(aq)$, $\text{CsOH}(aq)$



Common Weak Bases:



Amines such as triethylamine, $(\text{CH}_3\text{CH}_2)_3\text{N}$

Amphoteric substances:

These substances that have both acidic and basic properties and, thus, can react with both acids and bases: e.g. HCO_3^- , HSO_4^- and H_2PO_4^- .

