

Chapter 2

- 36) fluorine – F; chlorine – Cl; bromine – Br; sulfur – S; oxygen – O; phosphorous – P
- 38) As – arsenic; I – iodine; Xe – xenon; He – helium; C – carbon; Si – silicon
- 40) promethium – Pm; technetium – Tc
- 42)
- | | |
|--------------------------------|--------------------------------|
| a) Six: Be, Mg, Ca, Sr, Ba, Ra | c) Four; Ni, Pd, Pt, Ds |
| b) Five; O, S, Se, Te, Po | d) Six; He, Ne, Ar, Kr, Xe, Rn |
- 46)
- | | |
|--------------------------|--------------------------|
| a) $^{17}_8\text{O}$ | e) $^{131}_{53}\text{I}$ |
| b) $^{37}_{17}\text{Cl}$ | f) ^7_3Li |
| c) $^{60}_{27}\text{Co}$ | |
| d) $^{57}_{26}\text{Fe}$ | |
- 62)
- | | |
|--------------------|------------------------|
| a) cesium fluoride | d) manganese(IV) oxide |
| b) lithium nitride | e) titanium(IV) oxide |
| c) silver sulfide | f) strontium phosphide |
- 66)
- | | |
|-------------------------|------------------------------|
| a) dinitrogen tetroxide | c) sulfur dioxide |
| b) iodine trichloride | d) diphosphorus pentasulfide |
- 68)
- | | |
|------------------------|----------------------|
| a) acetic acid | g) sulfuric acid |
| b) ammonium nitrite | h) strontium nitride |
| c) cobalt(III) sulfide | i) aluminum sulfite |
| d) iodine monochloride | j) tin(IV) oxide |
| e) lead(II) phosphate | k) sodium chromate |
| f) potassium iodate | l) hypochlorous acid |
- 70)
- | | |
|---------------------------------|--|
| a) SF_2 | g) $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ |
| b) SF_6 | h) NH_4HSO_4 |
| c) NaH_2PO_4 | i) $\text{Co}(\text{NO}_3)_3$ |
| d) Li_3N | j) Hg_2Cl_2 |
| e) $\text{Cr}_2(\text{CO}_3)_3$ | k) KClO_3 |
| f) SnF_2 | l) NaH |
- 72)
- | | |
|----------------------------------|---------------------|
| a) $(\text{NH}_4)_2\text{HPO}_4$ | g) HBr |
| b) Hg_2S | h) HBrO_2 |
| c) SiO_2 | i) HBrO_4 |
| d) Na_2SO_3 | j) KHS |
| e) $\text{Al}(\text{HSO}_4)_3$ | k) CaI_2 |
| f) NCl_3 | l) CsClO_4 |